SNC2D Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Identifying Reactions Lab**

Purpose: To identify the types of reactions & to predict the products of the reactions.

Procedure: Obtain a test tube rack with 3 test tubes from the front of the classroom and take it to your lab bench.

**Part 1: Magnesium metal + Oxygen [teacher demonstration]**

1. Observe the burning of magnesium metal (Mg). \*Do not look directly at the flame.\*

2. Write your observations in the table.

**Part 2: Hydrogen peroxide**

1. Pour hydrogen peroxide (H2O2) into a test tube to a depth of 1 cm.

2. Add a small amount of yeast to the test tube.

*\*The yeast is a catalyst (makes the reaction go faster), it is not a reactant.\**

3. Tilt the test tube so that the yeast combines with the hydrogen peroxide.

4. Observe the test tube after 10 minutes. Write your observations in the table.

5. Dispose of the products in the sink.

**Part 3: Magnesium + Hydrochloric Acid**

1. Place a small piece of magnesium metal (Mg) in a test tube.

2. Add 10 drops of hydrochloric acid (HCl) to the test tube.

3. Observe the reaction and write your observations in the table.

4. Dispose of the products in the waste container provided.

**Part 4: Silver Nitrate + Sodium Chloride**

1. Place 2 drops of silver nitrate (AgNO3) in a test tube.

2. Add a dropper of sodium chloride (NaCl) to the test tube, one drop at a time.

3. Observe the reaction and write your observations in the table.

4. Dispose of the products in the waste container provided.

**Please wash all of your test tubes & put them upside down in the test tube rack.**

**Leave the test tube rack at the front of the classroom.**

SNC2D

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|  | ***Predict the products & balance the equation:***  **Reactants 🡪 Products** | ***Identify the type of reaction:*** |
| **Part 1** | Mg + O2 🡪 |  |
| Qualitative Observations*: (Describe the physical properties of the reactants, and what happened during the reaction.)* | |
| **Part 2** | H2O2 🡪 |  |
| Qualitative Observations: (Describe the physical properties of the reactants, and what happened during the reaction.) | |
| **Part 3** | Mg + HCl 🡪 |  |
| Qualitative Observations: (Describe the physical properties of the reactants, and what happened during the reaction.) | |
| **Part 4** | AgNO3 + NaCl 🡪 |  |
| Qualitative Observations: (Describe the physical properties of the reactants, and what happened during the reaction.) | |